THE

BRITISH JOURNAL OF NUTRITION

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References. At the end of the paper references should be given in alphabetical order according to the name of the first author of the publication quoted, names with prefixes being entered under the prefix, and should include the authors' initials; the title of the paper should not be included. Titles of journals should be abbreviated in accordance with the system used in the World List of Scientific Periodicals (1934: 2nd ed. Oxford University Press). Examples of such abbreviations will be found in the current numbers of the British Journal of Nutrition and useful lists have recently been published in the Journal of Physiology (1945, 104, 232) and by the Biological Council (A List of Abbreviations of the Titles of Biological Journals, obtainable from H. K. Lewis & Co. Ltd., 136 Gower Street, London, W.C.1). References to books and monographs should include the town of publication and the name of the publisher, as well as the date of publication and the number of the edition to which reference is made. Thus:

- Barnett, J. W. & Robinson, F. A. (1942). Biochem. J. 36, 364.
- Culbertson, C. C. & Thomas, B. H. (1934). Rep. Ia agric. Exp. Sta. 1933-4, p. 51.
- Doisy, E. A., Somogyi, M. & Shaffer, P. A. (1923). J. biol. Chem. 55, xxxi.
- Fairley, N. H. (1938). Nature, Lond., 142, 1156.
- Hennessy, D. J. (1941). Industr. Engng Chem. (Anal. ed.), 13, 216.
- King, H. (1941). J. chem. Soc. p. 338.
- Osborne, T. B. & Mendel, L. B. (1914*a*). J. biol. Chem. 17, 325.
- Osborne, T. B. & Mendel, L. B. (1914b). J. biol. Chem. 18, 1.
- Osborne, T. B. & Mendel, L. B. (1916). Biochem. J. 10, 534.
- Osborne, T. B., Mendel, L. B. & Ferry, E. L. (1919). J. biol. Chem. 37, 233.
- Starling, E. H. (1915). Principles of Human Physiology, 2nd ed. London: Churchill.

Statistical Treatment of Data. In general the publication is not necessary of all the individual results of a number of replicated tests. A statement of the number of individual results, their mean value, and some appropriate measure of their variability, is usually sufficient.

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Description of Solutions. Solutions of common $\operatorname{acids}_{\operatorname{Ei}}^{\operatorname{Ei}}$ bases and salts are preferably defined in terms of normality (N) or molarity (M), e.g. N-HCl; 0·1 M-NaH₂PO₄. The term is '%' must be used in its correct sense, i.e. g./100 g. of solution. 10% HCl means 10 g. of hydrogen chloride in the term is constant of a solution, and should never be used to be used.

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